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A chemical reaction rearranges the constituent atoms
of the reactants to create different substances as

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products. The properties of the products are different from those of the reactants. Chemical reactions differ from physical changes, which include changes of state, such as ice melting to water and water evaporating to vapor. If a physical change occurs, the physical properties of a substance will change, but its chemical identity will remain the same.

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Solution: A Because ascorbic acid acts as a monoprotic acid, we can write the balanced chemical equation for the reaction as $\text{[HAsc(aq) + OH}^{-}\text{(aq) \rightarrow Asc}^{-}\text{(aq) + H}_2\text{O(l)]}$ where HAsc is ascorbic acid and Asc⁻ is ascorbate. The number of moles of OH⁻ ions needed to neutralize the ascorbic acid is

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$$dy_j = \sum_{k=1}^N \nu_{kj} r_k dt$$

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Gibbs energy G versus reaction extent ϵ . Figure 3.4 (page 84): Partial pressures of components A and C versus liquid-phase composition in a nonideal solution. Figure 3.6 (page 95): Gibbs energy contours for the pentane reactions as a function of the two reaction extents.

4.6: Solution Stoichiometry and Chemical Analysis ...

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Ansys engineering simulation for reactor design including CFD for chemical reactor design, structural and thermal analysis and instrumentation and control software used for design and analysis of chemical reactors enable reaction engineers to answer what-if questions as they design and enhance reactors performance, energy usage, reactor yield and product uniformity.

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Solution The general solution method is best explained in terms of the complex reaction network notation where the rate for the j th species is given : $N_k j_1 y_1 k j_2 y_2 \dots k l_j y_j \dots dt l dy_j (1)$ A solution is assumed of the form: $y_j Y_j e (2)$ where Y_j and are constants to be determined.

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